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Molar Equations

The only way to accurately compare reactants and/or products is to compare the moles of each. In most situations you have a known mass or volume/concentration of a substance and you must convert to moles before you can compare them. There are three equations you can use to convert to moles.

Solids:

$$\text{Moles} = \frac{\text{Mass(g)}}{\text{Molar Mass(g/mol)}}$$

The molar mass of an element or compound can be found using the periodic table and has the units grams/mole. For this equation the mass needs to be in grams as well, to convert from milligrams to grams $\div 1000$ and to convert from kilograms to grams $\times 1000$.

Liquids:

$$\text{Moles} = \text{Volume (unit of volume)} \times \text{Concentration (moles/unit of volume)}$$

For this equation the units for volume and concentration must be consistent. For example: volume measured in dm^3 and concentration in moles/dm^3 you can immediately multiply. If the units for volume are not the same [volume in cm^3 and concentration in mol/dm^3] then you must change the units of your volume to match the concentration before you multiply.

Gases:

Molar volume is the amount of space a gas takes up at a given temperature and pressure. There are two molar gas constants, one at standard temperature and pressure (STP) and one at room temperature and pressure (RTP). The commonly used units for molar volume are dm^3 or Liters.

$$\text{Gases @ S.T.P.} \quad \text{Moles} = \frac{\text{Volume}}{22.4\text{dm}^3} \quad \text{OR} \quad \text{Moles} = \frac{\text{Volume}}{22.4\text{L}}$$

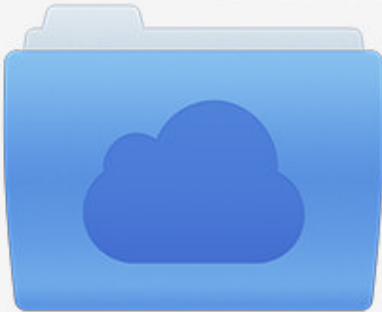
0°C and I.A.T.M.

$$\text{Gases @ R.T.P.} \quad \text{Moles} = \frac{\text{Volume}}{24\text{dm}^3} \quad \text{OR} \quad \text{Moles} = \frac{\text{Volume}}{24\text{L}}$$

25°C and I.A.T.M.

$$\begin{aligned} 1 \text{ cm}^3 &= 1 \text{ ml} \\ 1 \text{ dm}^3 &= 1 \text{ L} \end{aligned}$$





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